

Key word	Definition
balanced symbol equation	In a balanced symbol equation, chemical formulae represent the reactants and products. The equation shows how atoms are rearranged, and gives the relative amounts of reactants and products.
chemical reaction	A change in which atoms are rearranged to create new substances.
combustion	A chemical reaction in which a substance reacts quickly with oxygen and gives out light and heat.
conservation of mass	In a chemical reaction, the total mass of reactants is equal to the total mass of products. This is conservation of mass. Mass is conserved in chemical reactions and in physical changes.
decomposition	A chemical reaction in which a compound breaks down to form simpler compounds and/or elements.
discrete	A variable that can only have whole-number values.
endothermic change	An endothermic change transfers energy from the surroundings.
exothermic change	An exothermic change transfers energy to the surroundings.
fossil fuel	A fuel made from the remains of animals and plants that died millions of years ago. Fossil fuels include coal, oil, and natural gas.
fuel	A material that burns to transfer useful energy.
hazard	A possible source of danger.
non-renewable	Some fuels are non-renewable. They form over millions of years, and will one day run out.
oxidation	A chemical reaction in which substances react with oxygen to form oxides.
physical change	A change that is reversible, in which new substances are not made. Examples of physical changes include changes of state, and dissolving.
product	A substance that is made in a chemical reaction.
reactant	A starting substance in a chemical reaction.

risk	The chance of damage or injury from hazard.
word equation	A way of representing a chemical reaction simply. The reactants are on the lefts of an arrow, and the products are on the right. The arrow means <i>reacts to make</i> .